## PERIODIC TABLE TRENDS

Step 1: Complete the squares for each	
element by adding the atomic number, name, and atomic mass.	P =
Write the atomic number at the top of the square.  Write the atomic name under the symbol.	Be N =
Write the atomic mass at the bottom of the square.	E =
Step 2: Determine the number of electrons, protons, and neutrons in each statement.	Bohr Diagram Dot Structure
Step 3: Create a Bohr diagram for each element.	Be
Step 4: Draw a Lewis Dot Structure for each element.	Metal Nonmetal Metalloid → Solid Gas

Step 5: Show whether the element is a metal, nonmetal, or metalloid by circling the correct response. Also determine if the element is a solid or gas by circling the correct choice.

Step 6: Use the following colors to shade in the square for each element. You should ONLY color in the small square in the upper left-hand corner and not the entire card.

Green – Li and Na	Pink – O and S	Blue – Be and Mg
Purple – F and Cl	Orange – B and Al	Red – C and Si
Brown – N and P	Yellow - He Ne and Ar	White - H

Step 7: Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.



Step 8: Put a title on your table.

Step 9: Make a key for each color with the name of the group for that column.

Step 10: Write the column number at the top of each group and the period number on the left side of the table. Be careful, column numbers are not in order from 1-18 and the transition metals, columns 3-12 are not on the table.

Step 11: Draw the red zigzag line that separates the metals from the nonmetals.

Step 12: Answer the questions on the worksheet using the information on your periodic table.

Copyright: T. Trimpe 2002 and M. Robinson 2007