

Molarity, Molality, and Mole Fraction Worksheet (2) - Chemistry

Name: _____

Molarity (M)	Molality (m)	Mole Fraction (X)
$\frac{\text{moles of solute}}{1 \text{ Liter of solution}}$	$\frac{\text{moles of solute}}{1 \text{ kilogram of solvent}}$	$\frac{\text{moles of component}}{\text{total moles of solution}}$

Solution A

	solution	solute	solvent
Grams of	1100 g sol	280.8 g NaCl	819.2 g H ₂ O
Moles of	50.31 mol sol	4.8 mol NaCl	45.11 mol H ₂ O
Liters of	1 L		0.8192 L H ₂ O

What is the molarity of this solution? _____ M NaCl

What is the molality of this solution? _____ m NaCl

What is the mole fraction of this solution? _____ X NaCl