binary ionic compounds (Homework)

1.
A simple ion with a +1 charge (for example, Na^+) results when an atom _____ electrons.

2.
Positive ions are called _____.
whereas negative ions are called _____.

3.
Simple negative ions formed from single atoms are given names that end with the letters _____.

4.
How many electrons are contained in each of the following ions?

(a) K^+ (b) Mn^{2+} (c) Co^{3+} (d) Co^{2+}

(e) Cr^{3+} (f) I^- (g) Fe^{3+} (h) P^{3-}

5.
For the following pairs of ions, use the concept that a chemical compound must have a net charge of zero to predict the formula of the simplest compound that the ions are most likely to form. (Type your answer using the format CO2 for CO_2.)

(a) Na^+ and C^{4-} (b) Sn^{4+} and N^{3-} (c) Fe^{3+} and P^{3-} (d) Sn^{4+} and C^{4-}

(e) Mg^{2+} and N^{3-} (f) Mg^{2+} and C^{4-} (g) Fe^{3+} and C^{4-} (h) Fe^{3+} and S^{2-}

6.
Write the correct formula for the ionic compound composed of potassium and iodide. (Type your answer using the format CH4 for CH_4.)

7.