

Worksheet 5: Writing Formulas for Ionic Compounds and Acids

Review Questions:

Name each of the following compounds:

1. SnCl₄ _____
2. SnCl₂ _____
3. BaO _____
4. HCl _____
5. KOH _____
6. NH₄NO₃ _____
7. NaHCO₃ _____
8. HgO _____
9. Zn(NO₂)₂ _____
10. H₃PO₄ _____
11. HF _____

Writing Formulas:

Fill in the following information. Use your periodic table for help.

This section includes elements from all parts of the periodic table. Remember that roman numerals indicate charges for those in the transition metals, lead, and tin. Some include polyatomic ions. Hint: if it ends in -ate or -ite then it is a polyatomic ion!!!

Number	Name	Cation (include symbol and charge)	Anion (include symbol and charge)	Formula
12	Potassium iodide			
13	Copper (I) fluoride			
14	Iron (III) iodide			
15	Manganese (IV) oxide			
16	Calcium sulfide			
17	Iron (III) oxide			
18	Silver nitrate			
19	Sodium sulfate			
20	Beryllium nitrate			
21	Ammonium nitrate			
22	Cesium bromide			
23	Aluminum sulfide			
24	Aluminum bromide			
25	Sodium phosphide			
26	Potassium sulfide			

WORKSHEET 5 CONTINUES ON THE NEXT PAGE #27-33

1. Hydrobromic acid _____
2. Hydrosulfuric acid _____