

## **Molarity Practice Worksheet**

*Find the molarity of the following solutions:*

- 1) 0.5 moles of sodium chloride is dissolved to make 0.05 liters of solution.
  
- 2) 0.5 grams of sodium chloride is dissolved to make 0.05 liters of solution.
  
- 3) 0.5 grams of sodium chloride is dissolved to make 0.05 mL of solution.
  
- 4) 734 grams of lithium sulfate are dissolved to make 2500 mL of solution.
  
- 5)  $6.7 \times 10^{-2}$  grams of  $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_4$  are dissolved to make 3.5 mL of solution.
  
- 6) I have two solutions. In the first solution, 1.0 moles of sodium chloride is dissolved to make 1.0 liters of solution. In the second one, 1.0 moles of sodium chloride is added to 1.0 liters of water. Is the molarity of each solution the same? Explain your answer.

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