

- 1) What is the pH of a 0.0235 M HCl solution?

- 2) What is the pOH of a 0.0235 M HCl solution?

- 3) What is the pH of a 6.50×10^{-3} M KOH solution? (Hint: this is a basic solution – concentration is of OH⁻)

- 4) A solution is created by measuring 3.60×10^{-3} moles of NaOH and 5.95×10^{-4} moles of HCl into a container and then water is added until the final volume is 1.00 L. What is the pH of this solution?

- 5) What is the pH of a 6.2×10^{-8} M NaOH solution? (Hint: this is a basic solution – concentration is of OH⁻)

- 6) A solution with a H⁺ concentration of 1.00×10^{-7} M is said to be neutral. Why?