

## Naming Binary and Polyatomic Ionic Compounds

Determine the name of the following ionic compounds.

Reminder: If the compound is composed of only 2 elements, only your periodic table is needed to name the compound. The first element's name will be the same as the name on the periodic table. The ending of the second element will need to be changed to *ide*. If there are 3 or more elements in the formula you will need to use your polyatomic ion sheet for part of the name.

1.  $\text{Na}_2\text{O}$  \_\_\_\_\_
2.  $\text{NaOH}$  \_\_\_\_\_
3.  $\text{AlCl}_3$  \_\_\_\_\_
4.  $\text{KCl}$  \_\_\_\_\_
5.  $\text{BaCl}_2$  \_\_\_\_\_
6.  $\text{SrF}_2$  \_\_\_\_\_
7.  $\text{Ca}(\text{OH})_2$  \_\_\_\_\_
8.  $\text{Be}_2\text{SO}_4$  \_\_\_\_\_
9.  $\text{MgBr}_2$  \_\_\_\_\_
10.  $\text{Al}_2\text{S}_3$  \_\_\_\_\_
11.  $\text{MgO}$  \_\_\_\_\_
12.  $\text{TiI}_4$  \_\_\_\_\_
13.  $\text{Li}_3\text{PO}_4$  \_\_\_\_\_
14.  $\text{K}_3\text{N}$  \_\_\_\_\_
15.  $\text{Ca}(\text{CH}_3\text{COO})_2$  \_\_\_\_\_
16.  $\text{Ga}(\text{NO}_2)_3$  \_\_\_\_\_
17.  $\text{CS}_2\text{SO}_3$  \_\_\_\_\_
18.  $\text{NH}_4\text{OH}$  \_\_\_\_\_
19.  $\text{Al}(\text{CN})_3$  \_\_\_\_\_
20.  $\text{Mg}_3\text{P}_2$  \_\_\_\_\_