

$$K_{eq} = \frac{\text{Moles/liter (vapor)}}{\text{Moles/liter (aqueous)}}$$

	10 ²	
Alkanes	10	CH ₃ — CH ₃
Alkenes		CH ₂ = CH ₂
Alkynes	1	CH = CH
Thiols	10 ⁻¹	C ₂ H ₅ — SH
Chlorides		C ₂ H ₅ — Cl
Thioethers, ethers		CH ₃ — S — CH ₃
		CH ₃ — O — CH ₃
Esters	10 ⁻²	$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3 - \text{C} - \text{O} - \text{CH}_3 \end{array}$
Aldehydes	10 ⁻³	$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3 - \text{C} - \text{H} \\ \\ \text{CH}_3 \end{array}$
Ketones		$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3 - \text{C} - \text{CN} \end{array}$
Nitriles		
Amines	10 ⁻⁴	C ₂ H ₅ — NH ₂
Alcohols		C ₂ H ₅ — OH
Water	10 ⁻⁵	H ₂ O
Acids		$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3 - \text{C} - \text{OH} \end{array}$
Phosphotriesters	10 ⁻⁶	$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3 - \text{C} - \text{N}(\text{CH}_3)_2 \end{array} \quad (\text{CH}_3\text{O})_3\text{P} = \text{O}$
Amides	10 ⁻⁷	$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3 - \text{C} - \text{NH}_2 \end{array}$
Diols		$\begin{array}{c} \text{CH}_2\text{OH} \\ \\ \text{CH}_2\text{OH} \end{array}$
Peptides		$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3 - \text{C} - \text{NH}(\text{CH}_3) \end{array}$