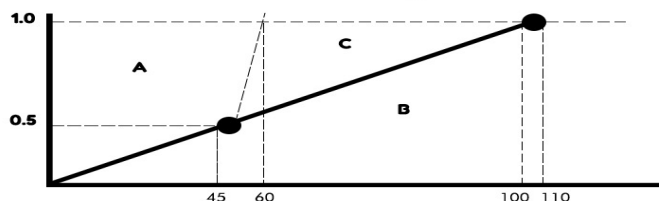


Name: _____

Date: _____

Phase Change Key



1. What is this substance's normal boiling point, at 1 atmosphere of pressure? **100°C**
2. Which section represents the solid phase? **A**
3. At a constant temperature, what would you do to cause this substance to change from the liquid phase to the solid phase? **You would need to increase the pressure.**
4. What does sublimation mean? **The process of going from a solid to a gas directly, skipping the liquid phase altogether.**
5. What is this substance's normal melting point, at 1 atmosphere of pressure? **60°C**
6. What section represents the liquid phase? **C**
7. What section represents the gas phase? **B**
8. What letter represents the triple point? **D**
9. Above what temperature is it impossible to liquefy this substance, no matter what the pressure? **110°C (This is known as the critical temperature)**
10. At what temperature and pressure do all three phases coexist? **45°C and 0.5 atm**