

Balancing Chemical Equations Worksheet

1. \_\_\_\_\_ H<sub>2</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ H<sub>2</sub>O
2. \_\_\_\_\_ N<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub> → \_\_\_\_\_ NH<sub>3</sub>
3. \_\_\_\_\_ S<sub>8</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ SO<sub>3</sub>
4. \_\_\_\_\_ N<sub>2</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ N<sub>2</sub>O
5. \_\_\_\_\_ HgO → \_\_\_\_\_ Hg + \_\_\_\_\_ O<sub>2</sub>
6. \_\_\_\_\_ CO<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O → \_\_\_\_\_ C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> + \_\_\_\_\_ O<sub>2</sub>
7. \_\_\_\_\_ Zn + \_\_\_\_\_ HCl → \_\_\_\_\_ ZnCl<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>
8. \_\_\_\_\_ SiCl<sub>4</sub> + \_\_\_\_\_ H<sub>2</sub>O → \_\_\_\_\_ H<sub>4</sub>SiO<sub>4</sub> + \_\_\_\_\_ HCl
9. \_\_\_\_\_ Na + \_\_\_\_\_ H<sub>2</sub>O → \_\_\_\_\_ NaOH + \_\_\_\_\_ H<sub>2</sub>
10. \_\_\_\_\_ H<sub>3</sub>PO<sub>4</sub> → \_\_\_\_\_ H<sub>4</sub>P<sub>2</sub>O<sub>7</sub> + \_\_\_\_\_ H<sub>2</sub>O
11. \_\_\_\_\_ C<sub>10</sub>H<sub>16</sub> + \_\_\_\_\_ Cl<sub>2</sub> → \_\_\_\_\_ C + \_\_\_\_\_ HCl
12. \_\_\_\_\_ CO<sub>2</sub> + \_\_\_\_\_ NH<sub>3</sub> → \_\_\_\_\_ OC(NH<sub>2</sub>)<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O
13. \_\_\_\_\_ Si<sub>2</sub>H<sub>3</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ SiO<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O<sub>3</sub>
14. \_\_\_\_\_ Al(OH)<sub>3</sub> + \_\_\_\_\_ H<sub>2</sub>SO<sub>4</sub> → \_\_\_\_\_ Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> + \_\_\_\_\_ H<sub>2</sub>O
15. \_\_\_\_\_ Fe + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ Fe<sub>2</sub>O<sub>3</sub>
16. \_\_\_\_\_ Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> + \_\_\_\_\_ KOH → \_\_\_\_\_ K<sub>2</sub>SO<sub>4</sub> + \_\_\_\_\_ Fe(OH)<sub>3</sub>
17. \_\_\_\_\_ C<sub>7</sub>H<sub>6</sub>O<sub>2</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ CO<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O
18. \_\_\_\_\_ H<sub>2</sub>SO<sub>4</sub> + \_\_\_\_\_ HI → \_\_\_\_\_ H<sub>2</sub>S + \_\_\_\_\_ I<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O
19. \_\_\_\_\_ FeS<sub>2</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ Fe<sub>2</sub>O<sub>3</sub> + \_\_\_\_\_ SO<sub>2</sub>
20. \_\_\_\_\_ Al + \_\_\_\_\_ FeO → \_\_\_\_\_ Al<sub>2</sub>O<sub>3</sub> + \_\_\_\_\_ Fe
21. \_\_\_\_\_ Fe<sub>2</sub>O<sub>3</sub> + \_\_\_\_\_ H<sub>2</sub> → \_\_\_\_\_ Fe + \_\_\_\_\_ H<sub>2</sub>O
22. \_\_\_\_\_ Na<sub>2</sub>CO<sub>3</sub> + \_\_\_\_\_ HCl → \_\_\_\_\_ NaCl + \_\_\_\_\_ H<sub>2</sub>O + \_\_\_\_\_ CO<sub>2</sub>
23. \_\_\_\_\_ K + \_\_\_\_\_ Br<sub>2</sub> → \_\_\_\_\_ KBr
24. \_\_\_\_\_ C<sub>7</sub>H<sub>16</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ CO<sub>2</sub> + \_\_\_\_\_ H<sub>2</sub>O
25. \_\_\_\_\_ P<sub>4</sub> + \_\_\_\_\_ O<sub>2</sub> → \_\_\_\_\_ P<sub>2</sub>O<sub>5</sub>