

Atomic Structure

Name: _____ Period: _____

1. The 3 particles of the atom are:

a. _____

b. _____

c. _____

Their respective charges are:

a. _____

b. _____

c. _____

2. The number of protons in one atom of an element determines the atom's _____, and the number of electrons determines _____ of an element.

3. The atomic number tells you the number of _____ in one atom of an element. It also tells you the number of _____ in a neutral atom of that element. The atomic number gives the "identity" of an element as well as its location on the

Periodic Table. No two different elements will have the _____ atomic number.

4. The _____ of an element is the average mass of an element's naturally occurring atom, or isotopes, taking into account the _____ of each isotope.

5. The _____ of an element is the total number of protons and neutrons in the _____ of the atom.

6. The mass number is used to calculate the number of _____ in one atom of an element. In order to calculate the number of neutrons you must subtract the _____ from the _____.

7. Give the symbol and number of protons in one atom of:

Lithium _____

Bromine _____

Iron _____

Copper _____

Oxygen _____

Mercury _____

Krypton _____

Helium _____