

# Periodic Trends

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Atomic Radius

1. What trend in atomic radius do you see as you go down a group/family on the periodic table?
2. What causes this trend?
3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
4. What causes this trend?
5. Circle the atom in each pair that has the largest atomic radius.
  - a. Al B c. S O e. Br Cl
  - b. Na Al d. O F f. Mg Ca
6. Put the following elements in order from smallest to largest atomic radius and explain why:  
C, O, Sn, Sr.
7. Define electronegativity
8. How does the ionic radius of a nonmetal compare with its atomic radius?
9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
10. What causes this trend?
11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
12. What causes this trend?
13. Circle the atom in each pair that has the greater electronegativity.
  - a. Ca Ga b. Li O c. Cl S d. Br As e. Ba Sr f. O S
14. Which group tends to form +1 ions?  
\_\_\_\_\_
15. Which group tends to form +2 ions?  
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