

Periodic Trends Worksheet

Name: _____ Date: _____



1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
3. Why does fluorine have a higher ionization energy than iodine?
4. Why do elements in the same family generally have similar properties?
5. Indicate whether the following properties increase or decrease from left to right across the periodic table.
6. What trend in atomic radius occurs across the periodic table? What causes this trend?
7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?
8. Circle the atom in each pair that has the largest radius.
 - a. Al or B
 - b. Na or Al
 - c. S or O
 - d. O or F
 - e. Br or Cl
 - f. Mg or Ca