

6. Compute the cost at 5.0 ¢ / kWh of operating an electric motor for 8.0 hours, which takes 15.0 amps at 110.0 volts.

$$\frac{15.0 \text{ C}}{1 \text{ sec}} \times \frac{110.0 \text{ J}}{\text{C}} = 1650 \frac{\text{J}}{\text{sec}} = 1650 \text{ watt} \times \frac{1 \text{ kWatt}}{1000 \text{ watt}} = 1.65 \text{ kWatt} \times 8.0 \text{ hr}$$

$$= 13.2 \text{ kWh} \times 5.0 \text{ ¢} = \boxed{0.66 \text{ ¢}}$$

7. How long would it take to deposit 100. grams of Al(s) from an electrolytic cell containing Al₂O₃ in solution at a current of 125 amps?

$$100. \text{ g Al} \times \frac{1 \text{ mol Al}}{26.98 \text{ g}} \times \frac{3 \text{ mol e}^-}{1 \text{ mol Al}} \times \frac{96500 \text{ C}}{1 \text{ mol e}^-} \times \frac{1 \text{ sec}}{125 \text{ C}} \times \frac{1 \text{ min}}{60 \text{ sec}} = \boxed{\frac{8580 \text{ sec}}{143 \text{ min}}}$$

8. Electrolysis of aqueous NaCl leads to the production of gaseous Cl₂. How much energy in kWh must have been used to produce 9.0718 x 10⁹ kg of Cl₂ when this type of cell operates at 4.6 volts and 3.0 x 10⁵ amps?

$$9.0718 \times 10^9 \text{ kg Cl}_2 \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mol Cl}_2}{70.90 \text{ g Cl}_2} \times \frac{2 \text{ mole}^-}{1 \text{ mol Cl}_2} \times \frac{96500 \text{ C}}{1 \text{ mole}^-} \times \frac{4.6 \text{ J}}{1 \text{ C}} \times \frac{1 \text{ kWh}}{3.6 \times 10^6 \text{ J}}$$

$$= \boxed{3.2 \times 10^{10} \text{ kWh}}$$

9. A current of 3.0 amps running for 45.0 minutes electrolyzes a solution containing a +2 metal ion. What is the identity of the metal if 2.67 grams of it plate out in this cell?

$$45.0 \text{ min} \times \frac{60 \text{ sec}}{1 \text{ min}} \times \frac{3.0 \text{ C}}{1 \text{ sec}} \times \frac{1 \text{ mole}^-}{96500 \text{ C}} \times \frac{1 \text{ mol X}}{2 \text{ mole}^-} = 0.0420 \text{ mol X}$$

$$\frac{2.67 \text{ g}}{0.0420 \text{ mol}} = \boxed{63.57 \text{ g/mol} = \text{Cu}}$$

10. Galvanizing is placing a thin layer of zinc on an iron nail to protect the nail from corrosion. Find the energy in kWh required to electrolyze 12.6 grams of zinc on an iron nail sitting in an electrolytic cell of Zn²⁺(aq) with a potential of 5.00 volts.

$$12.6 \text{ g Zn} \times \frac{1 \text{ mol Zn}}{65.39 \text{ g Zn}} \times \frac{2 \text{ mole}^-}{1 \text{ mol Zn}} \times \frac{96500 \text{ C}}{1 \text{ mole}^-} \times \frac{5.00 \text{ J}}{1 \text{ C}} \times \frac{1 \text{ kWh}}{3.6 \times 10^6 \text{ J}} = \boxed{0.0517 \text{ kWh}}$$