

Name _____ Period _____ Date _____

Building Molecules - Covalent Bonds

(Complete the chart for each element)

Element	Lewis structure	# of single electrons	# bonds the atom can form
Hydrogen			
Oxygen			
Chlorine			
Sulfur			
Carbon			
Nitrogen			
Iodine			

- Write the Lewis Structure for each element.
Find the total # of electrons you have to work with.
- Figure out the number of bonds needed.
- Write the symbol for the atom that can form the most bonds in the center and arrange the other atom symbols around it.
- Draw in your bonds and valence electrons.
Use one line for each pair of electrons that is shared.
- Double-check to make sure each atom has 8 valence electrons around it (except H).
Remember shared electrons count for each atom it is shared with.