

Introduction to Chemical Equilibrium

A reaction with a complete forward

direction reaches a state of equilibrium (products formed equal reacting reactants)

A reaction, regardless of the rate of complete forward

direction, reaches a state of equilibrium when the relative number of particles (molecules, atoms) remains constant

Directions: Several of the following are statements of chemical equilibrium. Check which are necessary conditions for such a complete reaction.

1. The reaction stops forever. _____

2. The reaction is reversible. _____

3. The rate of reverse reaction _____

4. The yield is zero. _____

5. Mass and energy both increase. _____

6. A small part has broken the rules. _____

7. Equilibrium. _____

8. Nothing happens. _____

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