

### Polyatomic Ions Chart

Formula	Name	Formula	Name
$\text{NH}_4^+$	Ammonium	$\text{CrO}_4^{-2}$	Chromate
$\text{NH}_3$	Ammonia	$\text{Cr}_2\text{O}_7^{-2}$	Dichromate
$\text{C}_2\text{H}_3\text{O}_2^-$	Acetate	$\text{MnO}_4^-$	Permanganate
$\text{CH}_3\text{COO}^-$	Acetate	$\text{MnO}_4^{-2}$	Manganate
$\text{CN}^-$	Cyanide	$\text{NO}_2^-$	Nitrite
$\text{CO}_3^{-2}$	Carbonate	$\text{NO}_3^-$	Nitrate
$\text{HCO}_3^-$	Bicarbonate	$\text{OH}^-$	Hydroxide
$\text{C}_2\text{O}_4^{-2}$	Oxalate	$\text{PO}_4^{-3}$	Phosphate
$\text{ClO}^-$	Hypochlorite	$\text{SCN}^-$	Thiocyanate
$\text{ClO}_2^-$	Chlorite	$\text{Fe}(\text{CN})_6^{-3}$	Ferricyanide
$\text{ClO}_3^-$	Chlorate	$\text{SO}_3^{-2}$	Sulfite
$\text{ClO}_4^-$	Perchlorate	$\text{SO}_4^{-2}$	Sulfate
$\text{S}_2\text{O}_3^{-2}$	Thiosulfate	$\text{HSO}_4^-$	Hydrogen sulfate
$\text{BrO}^-$	Hypobromite	$\text{IO}_3^-$	Iodate
$\text{AsO}_2^{-3}$	Arsenite	$\text{SeO}_4^{-2}$	Selenate
$\text{BrO}_3^-$	Bromate	$\text{HSO}_3^-$	Hydrogen sulfite

### Commonly Used Multivalent Metals

Symbol	Name	-ous	-ic
$\text{Fe}$ ( <i>Ferrum</i> )	Iron	+2	+3
$\text{Pb}$ ( <i>Plumbum</i> )	Lead	+2	+4
$\text{Sn}$ ( <i>Stannum</i> )	Tin	+2	+4
$\text{Hg}$ ( <i>Hydrargyrum</i> )	Mercury	+1	+2
$\text{Cu}$ ( <i>Cuprum</i> )	Copper	+1	+2

### Prefixes

Mono = one	Hexa = six
Di = two	Hepta = seven
Tri = three	Octa = eight
Tetra = four	Nona = nine
Penta = five	Deca = ten