Name: _____ **Balancing Act**

This activity will help you become skilled at balancing chemical equations.

1. Cut out the squares that the bottom of the paper.
2. Decide whether or not each equation is balanced. If it is not, decide how to balance it, putting the small squares as needed into the spaces.
3. Some spaces will not require a number.

1. ____ Zn + ____ O_2 → ____ ZnO
2. ____ N_2 + ____ H_2 → ____ NH_3
3. ____ B + ____ F_2 → ____ BF_3
4. ____ H_2 + ____ O_2 → ____ H_2O
5. ____ Fe + ____ H_2O → ____ Fe_2O_3 + ____ H_2
6. ____ HgO + ____ Cl_2 → ____ $HgCl$ + ____ O_2
7. ____ $KClO_3$ → ____ KCl + ____ O_2
8. ____ Cl_2 + ____ $NaBr$ → ____ $NaCl$ + ____ Br_2
9. ____ Na_2O + ____ CO_2 → ____ Na_2CO_3
10. ____ Al_2O_3 → ____ Al + ____ O_2

2	6	2	4	2	6	2	2	2	2
2	2	6	4	2	6	2	2	2	4
6	2								