

NUMBER OF ATOMS IN A FORMULA**ANSWERS**

Determine the number of atoms in the following chemical formulas.

1. 2NaCl $\text{Na} = 2$ $\text{Cl} = 2$
2. H_2SO_4 $\text{H} = 2$ $\text{S} = 1$ $\text{O} = 4$
3. 3KNO_3 $\text{K} = 3$ $\text{N} = 3$ $\text{O} = 9$
4. CaCl_2 $\text{Ca} = 1$ $\text{Cl} = 2$
5. $4\text{C}_2\text{H}_6$ $\text{C} = 8$ $\text{H} = 24$
6. $2\text{Ba}(\text{OH})_2$ $\text{Ba} = 2$ $\text{O} = 4$ $\text{H} = 4$
7. $3\text{NH}_4\text{Br}$ $\text{N} = 3$ $\text{H} = 12$ $\text{Br} = 3$
8. $4\text{Ca}_3(\text{PO}_4)_2$ $\text{Ca} = 12$ $\text{P} = 8$ $\text{O} = 32$
9. $2\text{Al}_2(\text{SO}_4)_3$ $\text{Al} = 4$ $\text{S} = 6$ $\text{O} = 24$
10. $3\text{Mg}(\text{NO}_3)_2$ $\text{Mg} = 3$ $\text{N} = 6$ $\text{O} = 18$
11. $6\text{Cu}(\text{NO}_3)_2$ $\text{Cu} = 6$ $\text{N} = 12$ $\text{O} = 36$
12. 4KMnO_4 $\text{K} = 4$ $\text{Mn} = 4$ $\text{O} = 16$
13. $2\text{H}_2\text{O}_2$ $\text{H} = 4$ $\text{O} = 4$
14. $3\text{H}_3\text{PO}_4$ $\text{H} = 9$ $\text{P} = 3$ $\text{O} = 12$
15. $2(\text{NH}_4)_3\text{PO}_4$ $\text{N} = 6$ $\text{H} = 24$ $\text{P} = 2$ $\text{O} = 8$
16. $2\text{Fe}_2\text{O}_3$ $\text{Fe} = 4$ $\text{O} = 6$
17. $\text{NaC}_2\text{H}_3\text{O}_2$ $\text{Na} = 1$ $\text{C} = 2$ $\text{H} = 3$ $\text{O} = 2$
18. $4\text{Mg}(\text{C}_2\text{H}_3\text{O}_2)_2$ $\text{Mg} = 4$ $\text{C} = 16$ $\text{H} = 24$ $\text{O} = 16$
19. $3\text{Hg}_2\text{Cl}_2$ $\text{Hg} = 6$ $\text{Cl} = 6$
20. $2\text{K}_2\text{SO}_3$ $\text{K} = 4$ $\text{S} = 2$ $\text{O} = 6$