

Hwa Chong Institution

Acids, Bases and Salts (Worksheet 2)

Name: _____ () Date: _____

Class: _____

Section A:

Answer **all** the questions in this section in the brackets provided

1. The table below gives information about three indicators:

Indicator	Colour at low pH	pH at which colour changes	Colour at high pH
Phenolphthalein	Colourless	8 – 10	Pink
Bromophenol Blue	Yellow	3 – 4.5	Purple
Methyl Red	Red	4.5 – 5	Green

What is the colour if each indicator was added separately to distilled water?

- | | Phenolphthalein | Bromophenol Blue | Methyl Red |
|----|------------------------|-------------------------|-------------------|
| A. | Colourless | Yellow | Green |
| B. | Colourless | Purple | Green |
| C. | Pink | Purple | Green |
| D. | Colourless | Yellow | Red |
- (B)

2. An element **Z** combusts in oxygen to give a white oxide, which is sparingly soluble in water. Upon addition of phenolphthalein, the solution turns pink. What is **Z** most likely to be?

- A. Sodium
B. Zinc
C. Phosphorus
D. Calcium
- (D)

3. Which of the following is commonly used to decrease the acidity in soil?

- A. Ammonium sulphate
B. Aqueous ammonia
C. Sodium hydroxide
D. Calcium hydroxide
- (D)

4. Which of the following will be least likely to react with aqueous sodium hydroxide?

- A. Carbon dioxide
B. Phosphorus pentoxide