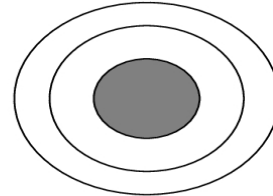


**Atomic Basics**

Name \_\_\_\_\_

**Part A: Atomic Structure**

1. Draw five protons in the nucleus of the atom. Label them with their charge.
2. Draw six neutrons in the nucleus of the atom.
3. Draw two electrons in the first energy level and label them with their charge.
4. Draw three electrons in the second energy level and label them with their charge.
5. What element is represented by the diagram? \_\_\_\_\_



**Part B: Atomic Calculations**

6. Label the information provided in the periodic table.

8	← _____
<b>O</b>	← _____
Oxygen	← _____
15.999	← _____

7. What does the atomic number represent?  
\_\_\_\_\_ or \_\_\_\_\_
8. What does the atomic mass represent?  
\_\_\_\_\_ + \_\_\_\_\_

9. How would you figure the number of protons or electrons in an atom?
10. How would you figure the number of neutrons in an atom?
11. Use your knowledge of atomic calculations to complete the chart.

Element	Atomic Number	Atomic Mass	Protons	Neutrons	Electrons
<b>Li</b>	3	7			
<b>P</b>	15	31			
<b>Cl</b>		35	17		
<b>Ni</b>	28			31	
<b>K</b>		39			19
<b>Ag</b>	47			61	
<b>H</b>		1	1		
<b>Si</b>				14	14
<b>W</b>			74	110	
<b>Ne</b>				10	10