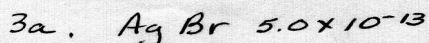
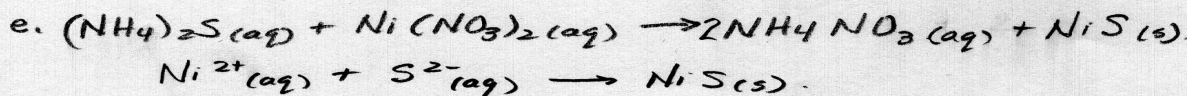
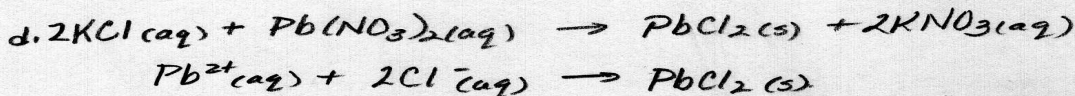
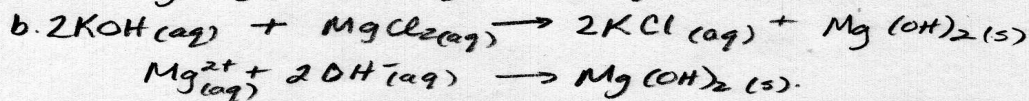
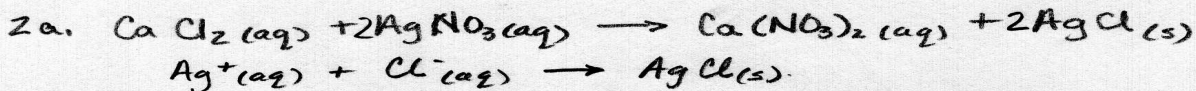


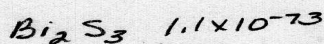
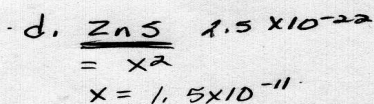
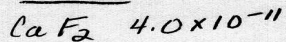
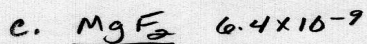
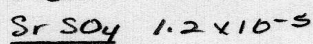
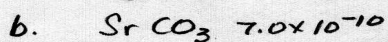
Solubility Equilibria worksheet.

- 1 a. $(\text{NH}_4)_2\text{S}$ soluble
 b. ZnCO_3 slightly soluble (I)
 c. $\text{Mg}(\text{OH})_2$ slightly soluble (I)
 d. Na_2CO_3 Soluble
 e. FeS slightly soluble (I)
 f. BaSO_4 insoluble
 g. Hg_2Cl_2 insoluble
 h. CaBr_2 Soluble
 i. CoS slightly soluble (I)

P1144

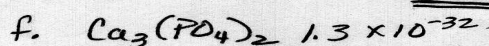
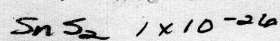


P A24



$= (2x)^2 \cdot (3x)^3 = 4x^2 \cdot 27x^3 = 108x^5$

$x = 1.0 \times 10^{-15}$



$= (3x)^3 (2x)^2$

$= (3x)^3 x$

$= 27x^3 \cdot 4x^2$

$= 27x^3 \cdot x$

$= 108x^5$

$= 27x^4$

$x = 4.1 \times 10^{-8}$

$x = 1.6 \times 10^{-5}$