Honors Chemistry Name_____

Reaction Prediction Worksheet Class _____ Date _____

Complete these problems ON YOUR OWN PAPER and attach your answers to this worksheet. For each reaction, you should:

a. state what type of reaction is expected
b. write the balanced equation for those reactions that do take place, and write symbols for the reactants for those reactions that do not occur
c. write net ionic equations for double replacement reactions

1. aluminum plus hydrochloric acid
2. calcium hydroxide plus nitric acid
3. aluminum plus magnesium
4. magnesium plus zinc nitrate
5. mercury plus oxygen (if mercury reacts, it will have a +1 charge)
6. zinc chloride plus hydrogen sulfide
7. dinitrogen pentoxide plus water
8. silver chloride plus sodium nitrate
9. sodium chlorate (heated)
10. barium nitrate plus sodium chromate
11. sodium bromide plus silver nitrate
12. calcium phosphate plus aluminum sulfate
13. zinc carbonate (heated)
14. mercury (I) sulfate plus ammonium nitrate
15. potassium plus fluorine
16. potassium nitrate plus zinc phosphate
17. lithium oxide plus water
18. sodium chloride (electrolyzed)
19. silver plus barium
20. iron (III) iodide plus copper (II) nitrate
21. copper plus sulfuric acid
22. lead plus potassium chlorate (if lead reacts, it will have a +2 charge)
23. sodium plus nitric acid
24. sulfur dioxide plus water
25. oxygen plus sulfur
26. sodium sulfate plus barium chloride
27. ammonium phosphate plus lithium hydroxide
28. mercury plus nitric acid (if mercury reacts, it will have a +2 charge)
29. sodium oxide plus water
30. calcium carbonate plus lithium chloride
31. iodine plus ammonium fluoride
32. ammonium sulfate plus barium hydroxide
33. barium oxide plus water
34. calcium plus oxygen
35. barium carbonate (heated)