

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Per.: \_\_\_\_\_

### Ionic Bonds Practice

1. Fill in the missing information on the chart.

Element	# of Protons	# of Electrons	# of Valence Electrons
Sodium			
Chlorine			
Beryllium			
Fluorine			
Lithium			
Oxygen			
Phosphorus			

2. For each of the following ionic bonds:

- Write the symbols for each element.
- Draw a Lewis Dot structure for the valence shell of each element.
- Draw an arrow (or more if needed) to show the transfer of electrons to the new element.
- Write the resulting chemical formula.
- Write the electron configurations for each ion that is formed. Ex.  $H^{1+} = 1s^2$

a) Sodium + Chlorine

b) Magnesium + Iodine

c) Sodium + Oxygen

d) Calcium + Chlorine

e) Aluminum + Chlorine