Unit 10b: Chemical Equations Nomenclature and Balancing Equations Practice Worksheet
I. For each chemical formula, list the number of each of the element(s) in it.
a. 3H ₂
b. 5H ₂ O
c. $4Al_2O_3$
d. 3S ₈
e. 7MgCl ₂
II. For each of the following formula equationsa. Balance the equations. Write the coefficients on the lines provided.b. Write the formula equation as a word equation
2 points each.
$\underline{\underline{\text{Example}}}_{\text{C}_4H_{10(g)}} \text{ (butane)} + \underline{\underline{\text{C}}_{2(g)}} \longrightarrow \underline{\underline{\text{C}}_{2(g)}} + \underline{\underline{\text{H}}_2O_{(g)}}$
a. $2C_4H_{10(g)}$ (butane) + $9O_{2(g)} \longrightarrow 4 CO_{2(g)} + 10H_2O_{(g)}$
b. Butane burns with oxygen gas to form carbon dioxide gas and water vapor.
1) $_$ Al + O ₂ \rightarrow $_$ Al ₂ O ₃
a.
b
2)NH ₃ +NH ₄ Cl
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Name___

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