WORKSHEET ON SINGLE & DOUBLE REPLACEMENT REACTIONS

Predict the products. Write formulas & balance each reaction. If there is no reaction, then just put NO RXN.

Single Replacement: A + BC \rightarrow B + AC or A + BC \rightarrow C + BA (when A and C are negative ions)

- 1. Zinc + Hydrogen chloride \rightarrow
- 2. Magnesium + Hydrogen Sulfate →
- 3. Copper (II) chloride + Flourine →
- 4. Silver + Sodium Hydroxide →
- 5. Potassium iodide + Bromine \rightarrow
- 6. Calcium + Hydrogen hydroxide →
- 7. Iron IV oxide + Hydrogen →

Double Replacement: $AB + CD \rightarrow AD + CB$

- 1. Barium chloride + Aluminum sulfate \rightarrow
- 2. Calcium nitride + water →
- 3. Calcium hydroxide + Hydrogen phosphate →
- 4. Hydrogen sulfate + Sodium hydrogen carbonate →
- 5. Calcium hydroxide + Ammonium chloride \rightarrow
- 6. Potassium iodide + Lead II Nitrate →
- 7. Sodium acetate + Calcium sulfide →

Complete each word equation, write formulas and balance the reaction equation. Then identify and place the type of reaction (single replacement or double replacement) in the blank provided.

- 1. Zinc + Silver nitrate \rightarrow
- 2. Aluminum + Hydrogen chloride →
- 3. Magnesium oxalate + Ammonium carbonate \rightarrow
- 4. Calcium + Aluminum nitrate →
- 5. Potassium flouride + Lead (II) Nitrate ightarrow
- 6. Calcium bromide + Silver nitrate ightarrow
- 7. Ammonium phosphate + Barium acetate \rightarrow
- 8. Sodium chloride + Potassium \rightarrow
- 9. Magnesium nitrate + ammonium chloride →
- 10. Iron (III) chlorate + calcium \rightarrow
- 11. Chlorine + Sodium bromide →
- 12. Potassium chloride + Silver nitrate →
- 13. Calcium hydroxide + Hydrogen nitrate ightarrow
- 14. Lead II nitrate + Potassium chloride →
- 15. Strontium carbonate + Hydrogen nitrate \rightarrow
- 16. Gold + Potassium nitrate →
- 17. Zinc + Silver nitrate →
- 18. Aluminum + Copper II sulfate →