

Predicting Chemical Equations Worksheet

	Complete the equation by predicting the products and balancing the equation	Type of Reaction
1	$\text{Al} + \text{S}_8 \Rightarrow$	
2	$\text{PbO} \Rightarrow$	
3	$\text{K} + \text{B}_2\text{O}_3 \Rightarrow$	
4	$\text{C} + \text{S}_8 \Rightarrow$	
5	$\text{HCl} + \text{NaOH} \Rightarrow$	
6	$\text{Na} + \text{NaNO}_3 \Rightarrow$	
7	$\text{Al}(\text{OH})_3 + \text{HBr} \Rightarrow$	
8	$\text{Na} + \text{O}_2 \Rightarrow$	
9	$\text{N}_2\text{O}_5 \Rightarrow$	
10	$\text{H}_3\text{PO}_4 + \text{KOH} \Rightarrow$	
11	$\text{NaOH} + \text{H}_2\text{CO}_3 \Rightarrow$	
12	$\text{KOH} + \text{HBr} \Rightarrow$	
13	$\text{H}_2\text{O}_2 \Rightarrow$	
14	$\text{Rb} + \text{RbNO}_3 \Rightarrow$	
15	$\text{C}_6\text{H}_6 + \text{O}_2 \Rightarrow$	
16	$\text{Al}(\text{OH})_3 + \text{H}_2\text{CO}_3 \Rightarrow$	
17	$\text{H}_2 + \text{O}_2 \Rightarrow$	
18	$\text{C}_2\text{H}_5\text{OH} + \text{O}_2 \Rightarrow$	
19	$\text{Cs} + \text{N}_2 \Rightarrow$	
20	$\text{Li} + \text{AlCl}_3 \Rightarrow$	
21	$\text{Mg} + \text{Cl}_2 \Rightarrow$	
22	$\text{C}_{10}\text{H}_{22} + \text{O}_2 \Rightarrow$	
23	$\text{NH}_3 \Rightarrow$	
24	$\text{C} + \text{SO}_2 \Rightarrow$	
25	$\text{C} + \text{O}_2 \Rightarrow$	
26	$\text{Al}(\text{OH})_3 + \text{H}_2\text{SO}_4 \Rightarrow$	
27	$\text{Na} + \text{O}_2 \Rightarrow$	
28	$\text{Rb} + \text{S}_8 \Rightarrow$	
29	$\text{Na} + \text{Cl}_2 \Rightarrow$	
30	$\text{C}_3\text{H}_8 + \text{O}_2 \Rightarrow$	
31	$\text{CaSO}_4 + \text{K}_3\text{PO}_4 \Rightarrow$	
32	$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \text{O}_2 \Rightarrow$	
33	$\text{Li} + \text{H}_2\text{O} \Rightarrow$	
34	$\text{NH}_4\text{OH} + \text{H}_3\text{PO}_4 \Rightarrow$	
35	$\text{Rb} + \text{P} \Rightarrow$	
36	$\text{CH}_4 + \text{O}_2 \Rightarrow$	
37	$\text{NH}_3 + \text{HCl} \Rightarrow$	
38	$\text{H}_3\text{PO}_4 + \text{Ca}(\text{OH})_2 \Rightarrow$	
39	$\text{NH}_3 + \text{O}_2 \Rightarrow$	
40	$\text{HgO} \Rightarrow$	