Name:					Date:
		Workshee	t # 3 (Chapte	er 13)	
solution prep		lving 10.0 mm		ne (nonvol	toult's Law predicts that attile) in 90.0 mmol
A) 0.498				0) 0.367	E) 0.0918
(torr) of wate nonelectroly	er above a soli te, MW = 180	ition prepared .0 g/mol) in 95	by dissolving .0 g of water	g 18.0 g of :?	
A) 24.3	B) 23.4	C) (D.451 D	0) 0.443	E) 23.8
(torr) of wate	er at 25 °C abo		prepared by c	dissolving	mine the vapor pressure 35 g of urea (a
A) 2.9	B) 3.3	C) 21	D) 27	E) (
depression co	onstant s 2.00 °C/m. V		zing point (°	C) of a so	olal freezing point lution prepared by of ethanol?
A) -115	B) -5.42	C) -132.3	D) -120.0		114.6
		int (°C) of a sc	dution prepar	red by dis	solving 11.3 g of
3.2			n 115 g of wa		molal freezing point
depression co		= 164 g/mol) i	n 115 g of wa		molal freezing point
depression co A) -3.34 6) A solutio point of -3.33	B) -1.11 n containing 1 3 °C. Given K g/mol.	= 164 g/mol) is ater is 1.86 °C/ C) 3.34 0.0 g of an unl t _f = 1.86°C/m	n 115 g of wa m. D) 1.11 known liquid for water, the	E) (and 90.0 molar mage	molal freezing point 0.00 g water has a freezing ass of the unknown liqui
depression co A) -3.34 6) A solutio point of -3.33	onstant for wa B) -1.11 n containing 1 3°C. Given K	= 164 g/mol) ir ater is 1.86 °C/ C) 3.34	n 115 g of wa m. D) 1.11 known liquid	E) (and 90.0	molal freezing point 0.00 g water has a freezing ass of the unknown liqui
depression of A) -3.34 6) A solutio point of -3.33 is A) 69.0 7) Calculate molal freezin	onstant for wa B) -1.11 In containing 1 3 °C. Given K g/mol. B) 333 the freezing p	= 164 g/mol) i ater is 1.86 °C/ C) 3.34 0.0 g of an unl c _f = 1.86°C/m	n 115 g of wa m. D) 1.11 known liquid for water, the D) 161 a 0.05500 m a of water is 1	E) (and 90.0 e molar max E) (aqueous so .86 °C/m.	molal freezing point 0.00 g water has a freezing ass of the unknown liqui
depression of A) -3.34 6) A solution point of -3.33 is A) 69.0 7) Calculate molal freezin A) 0.0286 8) Calculate molal freezin molal freezin A) 0.0286	onstant for war B) -1.11 n containing 13 °C. Given K g/mol. B) 333 the freezing pag-point-depre B) 0.1023 the freezing pag-point-depre Big-point-depre Big-point-depre	= 164 g/mol) inter is 1.86 °C/C) 3.34 0.0 g of an unlunction of the control of t	n 115 g of wa fm. D) 1.11 known liquid for water, the D) 161 a 0.05500 m is of water is 1 7 D) -0.102 a 0.05500 m is of water is 1	E) (and 90.0 e molar m E) (aqueous s .86 °C/m. 23 E) - aqueous s .86 °C/m.	molal freezing point 0.00 g water has a freezing ass of the unknown liqui- 62.1 colution of glucose. The 0.2046 colution of NaNO ₃ . The
depression of A) -3.34 6) A solutio point of -3.3; is A) 69.0 7) Calculate molal freezin A) 0.0286 8) Calculate	onstant for war B) -1.11 n containing 1 3 °C. Given K g/mol. B) 333 the freezing pag-point-depre B) 0.1023 the freezing pag-point-depre B) 0.1023	= 164 g/mol) inter is 1.86 °C/C) 3.34 0.0 g of an unlunction of the control of t	n 115 g of wa fm. D) 1.11 known liquid for water, the D) 161 a 0.05500 m is of water is 1 7 D) -0.102 a 0.05500 m is of water is 1	E) (and 90.0 e molar max E) (aqueous so .86 °C/m. 23 E) - aqueous so	molal freezing point 0.00 g water has a freezing ass of the unknown liqui- 62.1 colution of glucose. The 0.2046 colution of NaNO ₃ . The
depression of A) -3.34 6) A solution point of -3.33 is A) 69.0 7) Calculate molal freezin A) 0.0286 8) Calculate molal freezin A) 0.0286 9) A 0.15 m the percent in	onstant for war B) -1.11 n containing 13 °C. Given K g/mol. B) 333 the freezing pag-point-depre B) 0.1023 the freezing pag-point-depre B) -0.1023 aqueous solutionization of the	= 164 g/mol) is ater is 1.86 °C/C) 3.34 0.0 g of an unlustrial of the control of	m 115 g of warm. D) 1.11 known liquid for water, the D) 161 a 0.05500 m a of water is 1 7 D) -0.102 a 0.05500 m a of water is 1 0.1023 I acid has a free t this concentration.	E) (land 90.0) e molar ma E) (land 90.0) e molar ma E) (land 90.0) e aqueous sa .86 °C/m. 23 E) - aqueous sa .86 °C/m. O) -0.0562	molal freezing point 0.00 g water has a freezing ass of the unknown liqui- 62.1 colution of glucose. The 0.2046 colution of NaNO ₃ . The