

Writing Compound Formulas Worksheet

Write the correct formula for each of the following ionic compound names.

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| 1. Copper (I) phosphate _____ | 16. Iron (III) sulfate _____ |
| 2. Magnesium nitrite _____ | 17. Silver (I) dichromate _____ |
| 3. Cobalt (II) hydroxide _____ | 18. Lead (II) nitride _____ |
| 4. Iron (III) chromate _____ | 19. Strontium hypochlorite _____ |
| 5. Chromium (II) chloride _____ | 20. Barium chlorate _____ |
| 6. Zinc (II) fluoride _____ | 21. Tin (IV) oxide _____ |
| 7. Mercury (II) acetate _____ | 22. Nickel (II) cyanide _____ |
| 8. Manganese (II) iodide _____ | 23. Aluminum sulfide _____ |
| 9. Copper (II) nitrate _____ | 24. Lead (IV) chromate _____ |
| 10. Aluminum hydroxide _____ | 25. Potassium permanganate _____ |
| 11. Iron (II) phosphate _____ | 26. Calcium dihydrogen phosphate _____ |
| 12. Magnesium sulfite _____ | 27. Barium hydrogen sulfate _____ |
| 13. Tin (II) bromide _____ | 28. Lithium monohydrogen phosphate _____ |
| 14. Sodium bicarbonate _____ | 29. Silver (I) bisulfite _____ |
| 15. Chromium (III) oxalate _____ | 30. Ammonium carbonate _____ |

Write the Compound formulas for the covalent molecules below:

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| 31. Dinitrogen oxide _____ | 35. Tetraphosphorus trisulfide _____ |
| 32. Carbon tetraiodide _____ | 36. Sulfur hexafluoride _____ |

ii	N ⁺	nitrate	NO ₃ ⁻	KNO ₃	potassium nitrate	potassium
	Ba ⁺²	bromide	Br ⁻	BaBr ₂	barium bromide	barium
n	Sr ⁺²	sulfate	SO ₄ ⁻²	SrSO ₄	strontium sulfate	strontium
	Fe ⁺³	oxide	O ⁻²	Fe ₂ O ₃	iron(III) oxide	iron(III)
	Ag ⁺¹	chloride	Cl ⁻	AgCl	silver(I) chloride	silver(I)
n	Al ⁺³	carbonate	CO ₃ ⁻²	Al ₂ (CO ₃) ₃	aluminum carbonate	aluminum
	Au ⁺²	nitride	N ⁻³	Au ₃ N ₂	gold(II) nitride	gold(II)
	Cs ⁺¹	phosphate	PO ₄ ⁻³	Cs ₃ PO ₄	cesium phosphate	cesium
m	NH ₄ ⁺	sulfide	S ⁻²	(NH ₄) ₂ S	ammonium sulfide	ammonium
	Zn ⁺²	nitrite	NO ₂ ⁻	Zn(NO ₂) ₂	zinc(II) nitrite	zinc(II)
(I)	Cr ⁺¹	nitride	N ⁻³	Cr ₃ N	chromium(I) nitride	chromium
II)	Hg ⁺²	oxide	O ⁻²	HgO	mercury(II) oxide	mercury(II)