## NAMING AND WRITING FORMULAS FOR BINARY COMPOUNDS

Example	es: silver chloride: aluminum oxide:	Ag <sup>1+</sup> and Cl <sup>1-</sup> Al <sup>3+</sup> and O <sup>2-</sup>	AgCl Al <sub>2</sub> O <sub>3</sub>
	barium chloride: barium oxide:	Ba <sup>2+</sup> and Cl <sup>1-</sup> Ba <sup>2+</sup> and O <sup>2-</sup>	BaCl <sub>2</sub> BaO
1. barium chloride		2. calcium oxide	
3. silver bromide		4. aluminum sulfide	
5. iron(III) sulfide		6. iron (II) sulfide	
7. sodium oxide			
B. Look for the pa	ntterns when the elemen	nts from different fa	ide (N <sup>3-</sup> )amilies (groups) react with each other! the alkali metals and the halogens.
B. Look for the position 1. Write formula	atterns when the elements to	nts from different for	unilies (groups) react with each other!
B. Look for the position of th	ntterns when the elements to for four compounds to the formula for four compounds to the formula for four compounds to the formula for	nts from different fa using elements from using elements from	umilies (groups) react with each other! the alkali metals and the halogens.
B. Look for the positive formula in the second seco	ntterns when the elements to some some some some some some some som	nts from different for using elements from using elements from using elements from	unilies (groups) react with each other! the alkali metals and the halogens. the alkali metals and group VI.