Ionic Formula to Ionic Names Practice Worksheet

Ionic compounds are created when non-metallic atoms borrow electrons from metallic atoms. The ions created attract each other electrostatically.

Naming Rule: Name the metallic element first followed by the non-metallic element that has it end removed and replaced by the suffix – ide.

Carbon→ carbide	nitrogen→ nitride	Oxygen→ oxide	fluorine→ fluoride
phosphorus→ phosphide	sulfur→ sulfide	chlorine→ chloride	arsenic→ arsenide
bromine→ bromide	iodide→iodide		

Note: The number subscripts are not included in the naming, however, they identify how many of each element belongs in the compound.

1. Name the following ionic compounds:

1) $CaCl_2$ _____
2) Mg_3N_2_____
3) $BeBr_2$ _____
4) Na_3P _____
5) LiCl _____
6) AgAs _____
7) MgO _____
8) AgBr _____
9) $Ca_3\ N_2$ _____
10) Al_4C_3 _____
11) HF _____
12) CaO _____
13) K_2S _____
14) Ag_3N _____
15) $Sr\ I_2$ _____

2. Underline the metallic elements in each of the formulas and names above.

3. To the right of each name, identify the number of atoms of each element found in the compound.

Ex. NaCl → sodium chloride → one atom of sodium, one atom of chlorine

ZnI_2 →zinc iodide → one atom of zinc, two atoms of iodine