

Physical Science Worksheet: History of the Periodic Table

Short Answer

1. How did chemists change Mendeleev's periodic table in the early 1900s?
2. What prediction did Mendeleev make that came true less than 20 years later?
3. Phosphorus-33 (atomic number 15) contains how many electrons, protons, and neutrons?
4. A(n) _____ is an atom, or bonded group of atoms, that has a positive or negative charge.
5. An atom becomes negatively charged by _____.
6. What particle has a positive charge?
7. Very energetic particles that move in all directions around the nucleus of an atom are
8. A negative ion is known as a(n)
9. What information in the periodic table indicates the number of protons in an atom?
10. According to Dalton's theory of atoms, all atoms in any element
11. Atoms have no electric charge because they
12. The first person who came up with the idea of atoms was
13. What did Dalton's theory of atoms say about compounds?
14. As the mass number of an element's isotopes of an element increases, the number of protons
15. Substances that CANNOT be broken down chemically into other substances are
16. A positive ion is known as a(n)
17. The charge of an electron is
18. How many electrons does an atom generally need in its outer level to be the most stable?
19. An aluminum isotope consists of 13 protons, 13 electrons, and 14 neutrons. Its mass number is
20. Chlorine has atomic number 17 and mass number 35. It has how many protons, electrons, and neutrons?
21. Isotopes are atoms of the same element that have different
22. Neon-22 contains 12 neutrons. It also contains how many protons?
23. The modern periodic table is arranged in order of increasing
24. Mendeleev created the first periodic table by arranging elements in order of
25. Electrons involved in bonding between atoms are