

Worksheet I - Binary Compounds and Binary Acids

1. $\text{Al}_2\text{O}_3$ <i>aluminum oxide</i>	sodium sulfide $\text{Na}_2\text{S}$
2. $\text{SiI}_4$ <i>silicon tetraiodide</i>	calcium bromide $\text{CaBr}_2$
3. $\text{AsF}_3$ <i>arsenic trifluoride</i>	barium sulfide $\text{BaS}$
4. $\text{Cs}_3\text{N}$ <i>cesium nitride</i>	phosphorus triiodide $\text{PI}_3$
5. $\text{SiO}_2$ <i>silicon dioxide</i>	sulfur trioxide $\text{SO}_3$
6. $\text{N}_2\text{O}_5$ <i>dinitrogen pentoxide</i>	aluminum selenide $\text{Al}_2\text{Se}_3$
7. $\text{LiCl}$ <i>lithium chloride</i>	silicon disulfide $\text{SiS}_2$
8. $\text{B}_2\text{S}_3$ <i>diboron trisulfide</i>	carbon dioxide $\text{CO}_2$
9. $\text{K}_2\text{O}$ <i>potassium oxide</i>	sodium oxide $\text{Na}_2\text{O}$
10. $\text{PF}_5$ <i>phosphorus pentafluoride</i>	carbon monoxide $\text{CO}$
11. $\text{BeO}$ <i>beryllium oxide</i>	lithium fluoride $\text{LiF}$
12. $\text{SeCl}_2$ <i>selenium dichloride</i>	boron trichloride $\text{BCl}_3$
13. $\text{I}_2\text{O}_7$ <i>diiodine heptaoxide</i>	lithium phosphide $\text{Li}_3\text{P}$
14. $\text{BaI}_2$ <i>barium iodide</i>	diarsenic pentoxide $\text{As}_2\text{O}_5$
15. $\text{HCl}(\text{aq})$ <i>hydrochloric acid</i>	aluminum oxide $\text{Al}_2\text{O}_3$
16. $\text{MgS}$ <i>magnesium sulfide</i>	hydrofluoric acid $\text{HF}(\text{aq})$
17. $\text{H}_2\text{S}(\text{aq})$ <i>hydrosulfuric acid</i>	strontium fluoride $\text{SrF}_2$
18. $\text{Na}_3\text{N}$ <i>sodium nitride</i>	potassium sulfide $\text{K}_2\text{S}$
19. $\text{OF}_2$ <i>oxygen difluoride</i>	cesium telluride $\text{Cs}_2\text{Te}$
20. $\text{HI}(\text{aq})$ <i>hydroiodic acid</i>	hydroselenic acid $\text{H}_2\text{Se}$