

Empirical and Molecular Worksheet

1. Write the Empirical Formula, Empirical Mass and Molecular Mass for each of the following:

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| a. P_4O_6 _____ | e. $C_6H_8O_6$ _____ |
| b. C_6H_9 _____ | f. $C_{10}H_{22}$ _____ |
| c. CH_2OHCH_2OH _____ | g. $Cu_2C_2O_4$ _____ |
| d. $BrCl_2$ _____ | h. Hg_2F_2 _____ |

2. Write the Empirical Formula for Each of the Following:

a. A compound composed of: 72% iron and 27.6% oxygen by mass.

b. A compound composed of: 9.93% carbon, 58.6% chlorine, and 31.4% fluorine.

c. A compound composed of: 5.56g carbon and .0933g hydrogen.

3. Write the Molecular Formula for Each of the Following:

a. A compound with a molecular mass of 70g and an empirical formula of CH_2 .

b. A compound with a molecular mass of 46.0g and an empirical formula of NO_2 .