

Naming Binary and Polyatomic Ionic Compounds

Determine the name of the following ionic compounds.

Reminder: If the compound is composed of only 2 elements, only your periodic table is needed to name the compound. The first element's name will be the same as the name on the periodic table. The ending of the second element will need to be changed to *ide*. If there are 3 or more elements in the formula you will need to use your polyatomic ion sheet for part of the name.

1. Na_2O _____
2. NaOH _____
3. AlCl_3 _____
4. KCl _____
5. BaCl_2 _____
6. SrF_2 _____
7. $\text{Ca}(\text{OH})_2$ _____
8. Be_2SO_4 _____
9. MgBr_2 _____
10. Al_2S_3 _____
11. MgO _____
12. TiI_4 _____
13. Li_3PO_4 _____
14. K_3N _____
15. $\text{Ca}(\text{CH}_3\text{COO})_2$ _____
16. $\text{Ga}(\text{NO}_2)_3$ _____
17. CS_2SO_3 _____
18. NH_4OH _____
19. $\text{Al}(\text{CN})_3$ _____
20. Mg_3P_2 _____