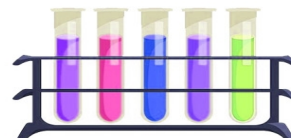


Name: \_\_\_\_\_

Date: \_\_\_\_\_

# Atomic and Mass Numbers



1. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminium
2. Why do elements in the same family generally have similar properties?
3. What trend in ionization energy occurs across a period on the periodic table?  
What causes this trend?
4. Why does fluorine have a higher ionization energy than iodine?
5. Rank the following elements by increasing atomic radius: carbon, aluminium.
6. What is the difference between electronegativity and ionization energy?
7. Why does fluorine have a higher ionization energy value than iodine?