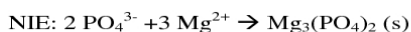
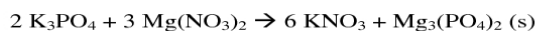


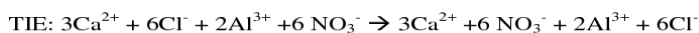
WRITING NET IONIC EQUATIONS

Write a balanced net ionic equation for the reaction that occurs in each of the following cases. Assume that all soluble reactants are added in the form of aqueous solutions. Indicate gases and precipitates that are formed, as well as insoluble solid reactants. If no reaction occurs, then write "NO RXN" and do not write a balanced equation.

1. potassium phosphate + magnesium nitrate \longrightarrow

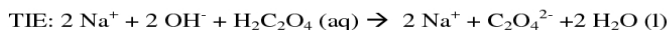
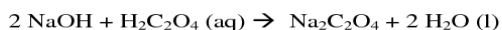


2. calcium chloride + aluminum nitrate \longrightarrow

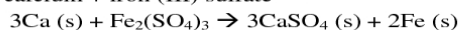


NIE: (no rxn)

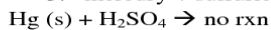
3. sodium hydroxide + oxalic acid \longrightarrow



4. calcium + iron (III) sulfate \longrightarrow



5. mercury + sulfuric acid \longrightarrow



TIE: no rxn

NIE: no rxn

6. calcium carbonate + sulfuric acid \longrightarrow

