

### Definitions

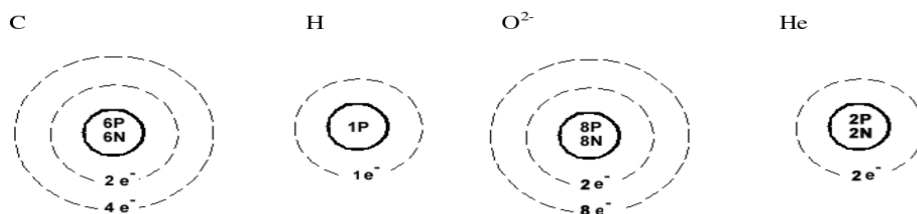
**Molecule**- a group of atoms held together by **covalent bonds** that displays all the properties of that compound. (Usually 2 or more non-metals in covalent bond(s))

**Compound**- a substance that is formed from different types of atoms and can be broken into its constituents by chemical reactions.

### Fill in the blank

Charged atoms form ionic bonds. Atoms that share electrons form covalent bonds. The two types of this bond are nonpolar covalent bonds and polar covalent bonds. Two polar molecules containing hydrogen may bond using hydrogen bonds.

**Draw the atomic structure of the following atoms:**



**Identify how many protons and electrons are in each element.**

	C	Be	Fe <sup>2+</sup>	Si	Cl
Neutrons	6	5	30	14	18
Protons	6	4	26	14	17
Electrons	6	4	24	14	18