

Name: Key  
 Hour: \_\_\_\_\_ Date: \_\_\_\_\_

**Chemistry: Graham's Law**

Do the following problems, showing your work and including all proper units.

1. If neon gas travels at 400 m/s at a given temperature, calculate the velocity of butane, C<sub>4</sub>H<sub>10</sub>, at the same temperature.

Ne = 20  
 C<sub>4</sub>H<sub>10</sub> = 58  

$$\sqrt{\frac{m_1}{m_2}} = \frac{v_2}{v_1}$$

$$\sqrt{\frac{20}{58}} = \frac{v_2}{400}$$

$$v_2 = 235 \text{ m/s}$$

2. Hydrogen sulfide, H<sub>2</sub>S, has a very strong rotten egg odor. Methyl salicylate, C<sub>8</sub>H<sub>8</sub>O<sub>3</sub>, has a wintergreen odor, and benzaldehyde, C<sub>7</sub>H<sub>6</sub>O, has a pleasant almond odor. If the vapors for these three substances were released at the same time from across a room, which odor would you smell first? Show your work and explain your answer.

H<sub>2</sub>S = 2(1) + 32 = 34 ← smallest mass ⇒ highest velocity  
 C<sub>8</sub>H<sub>8</sub>O<sub>3</sub> = 8(12) + 8(1) + 3(16) = 152  
 C<sub>7</sub>H<sub>6</sub>O = 7(12) + 6(1) + 16 = 106

3. A nitrogen molecule travels at about 505 m/s at room temperature. Find the velocity of a helium atom at the same temperature.

N<sub>2</sub> = 28 v = 505 m/s  
 He = 4  

$$\sqrt{\frac{m_1}{m_2}} = \frac{v_2}{v_1}$$

$$\sqrt{\frac{28}{4}} = \frac{v_2}{505}$$

$$v_2 = 1336 \text{ m/s}$$

4. A carbon dioxide molecule travels at 45.0 m/s at a certain temperature. At the same temperature, find the velocity of an oxygen molecule.

CO<sub>2</sub> = 44 v = 45.0 m/s  
 O<sub>2</sub> = 32  

$$\sqrt{\frac{m_1}{m_2}} = \frac{v_2}{v_1}$$

$$\sqrt{\frac{44}{32}} = \frac{v_2}{45.0}$$

$$v_2 = 52.8 \text{ m/s}$$

5. Nitrogen gas effuses through an opening 1.59 times faster than does an unknown gas.

a. Calculate the molecular mass of the unknown gas.

N<sub>2</sub> = 28  

$$\sqrt{\frac{m_1}{m_2}} = \frac{v_2}{v_1} = 1.59 = \sqrt{\frac{m_1}{28}}$$

$$2.53 = \frac{m_1}{28}$$

$$m_1 = 71$$

b. Make a reasonable prediction as to what the unknown gas is.

Cl<sub>2</sub>? = 70.9 amu

6. An unknown gas diffuses 1.62 times slower than does oxygen gas.

a. Calculate the molecular mass of the unknown gas.

O<sub>2</sub> = 32  

$$\sqrt{\frac{m_1}{m_2}} = \frac{v_2}{v_1} = 1.62 = \sqrt{\frac{m_1}{32}}$$

$$2.62 = \frac{m_1}{32}$$

$$m_1 = 84$$

b. Make a reasonable prediction as to what the unknown gas is. Kr = 83.8 amu

- Answers: 1. 236 m/s 4. 52.8 m/s 6a. 84 a.m.u.  
 2. H<sub>2</sub>S 5a. 71 a.m.u. 6b. Kr  
 3. 1336 m/s 5b. Cl<sub>2</sub>