

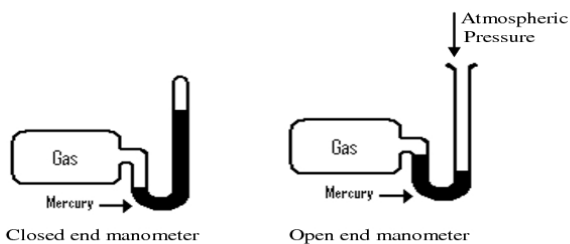
Chem Gas Worksheet #1. Blk ____ Name _____

Data to know & use! $1\text{ atm} = 760.\text{ mmHg} = 101.325\text{ kPa} = 14.7\text{ lb/in}^2$, or by 1989 definition
 Standard Pressure = $100.\text{ kPa} = 750.1\text{ mmHg} = 14.5\text{ lb/in}^2 = 0.987\text{ atm} = 1\text{ Bar}$.
 1 mole gas @ STP = $22.4\text{ L} = 22,400\text{ cm}^3$, STP = 0°C , 1 atm , $0\text{ K} = -273.15^\circ\text{C} = -459.67^\circ\text{F}$.
 1 mole gas @ SATP = $24.8\text{ L} = 24,800\text{ cm}^3$, Standard Ambient Temp. & Pressure = $100.\text{ kPa}$, 25°C .

A. Pressures. Show a unit cancellation setup. WATCH SIG. FIGS.

- $412\text{ mmHg} = \underline{\hspace{2cm}}\text{ atm}$.
- $760.\text{ kPa} = \underline{\hspace{2cm}}\text{ mmHg}$
- $14.7\text{ atm} = \underline{\hspace{2cm}}\text{ kPa}$
- $101.325\text{ lb/in}^2 = \underline{\hspace{2cm}}\text{ kPa}$
- $22.4\text{ mmHg} = \underline{\hspace{2cm}}\text{ kPa}$

B. Manometers.



- In a closed end manometer, the mercury level was 690. mm higher on the closed end than on the gas side. What was the pressure of the gas?
 $\underline{\hspace{2cm}}\text{ mmHg}$
- In a closed end manometer, the Hg levels were 419 mm different. What was the gas pressure?
 $\underline{\hspace{2cm}}\text{ mmHg}$
- In a closed end manometer, the Hg levels were 1273 mm different. What was the gas pressure IN ATM?
 $\underline{\hspace{2cm}}\text{ atm}$
- Open end manometer: atmospheric pressure 760. mmHg, and the mercury level was 120. mm higher on the right side than the left. What was the gas pressure?
 $\underline{\hspace{2cm}}\text{ mmHg}$
- Open end manometer, atmospheric pressure 755 mmHg, Hg level 75 mm higher on the left. What was the gas pressure?
 $\underline{\hspace{2cm}}\text{ mmHg}$
- Open end manometer, with the atmospheric pressure 97.2 kPa. Mercury level 35 mm higher on the left. What is the gas pressure?
 $\underline{\hspace{2cm}}\text{ kPa}$

C. Temperatures. a. $25^\circ\text{C} = \underline{\hspace{2cm}}\text{ K}$ b. $-147^\circ\text{C} = \underline{\hspace{2cm}}\text{ K}$ c. $926\text{ K} = \underline{\hspace{2cm}}^\circ\text{C}$

- d. $35.2\text{ K} = \underline{\hspace{2cm}}^\circ\text{C}$ e. $-2.8^\circ\text{C} = \underline{\hspace{2cm}}\text{ K}$ f. $12,780,000\text{ K} = \underline{\hspace{2cm}}^\circ\text{C}$