

Packet for Test 4

Worksheet Balancing Equations Chapter 10

The number at the end of the equation is the sum of the coefficients used to balance the equation. Understood ones are not included in the sum of the coefficients.

- _____1. $\text{Zn} + \text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$ (2)
- _____2. $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$ (0)
- _____3. $\text{Al} + \text{HCl} \rightarrow \text{AlCl}_3 + \text{H}_2$ (13)
- _____4. $\text{Mg} + \text{H}_3\text{PO}_4 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + \text{H}_2$ (8)
- _____5. $\text{Cu} + \text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + \text{Ag}$ (4)
- _____6. $\text{Ca} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{Pb} + \text{Ca}(\text{NO}_3)_2$ (0)
- _____7. $\text{Al} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{Pb} + \text{Al}(\text{NO}_3)_3$ (10)
- _____8. $\text{Zn} + \text{Sn}(\text{NO}_3)_4 \rightarrow \text{Zn}(\text{NO}_3)_2 + \text{Sn}$ (4)
- _____9. $\text{Cl}_2 + \text{AlI}_3 \rightarrow \text{AlCl}_3 + \text{I}_2$ (10)
- _____10. $\text{Br}_2 + \text{CuI} \rightarrow \text{CuBr} + \text{I}_2$ (4)
- _____11. $\text{NH}_4\text{OH} + \text{FeCl}_3 \rightarrow \text{NH}_4\text{Cl} + \text{Fe}(\text{OH})_3$ (6)
- _____12. $\text{KBr} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{KNO}_3 + \text{PbBr}_2$ (4)
- _____13. $\text{AlCl}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{HCl}$ (11)
- _____14. $\text{Al}_2(\text{SO}_4)_3 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + \text{AlCl}_3$ (8)
- _____15. $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + \text{NaCl}$ (2)
- _____16. $\text{H}_2\text{SO}_4 + \text{AlBr}_3 \rightarrow \text{HBr} + \text{Al}_2(\text{SO}_4)_3$ (11)
- _____17. $\text{Na}_3\text{PO}_4 + \text{BaCl}_2 \rightarrow \text{Ba}_3(\text{PO}_4)_2 + \text{NaCl}$ (11)
- _____18. $\text{K}_3\text{PO}_4 + \text{MgSO}_4 \rightarrow \text{Mg}_3(\text{PO}_4)_2 + \text{K}_2\text{SO}_4$ (8)
- _____19. $\text{NaOH} + \text{CuSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{Cu}(\text{OH})_2$ (2)
- _____20. $\text{Mg}(\text{OH})_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + \text{H}_2\text{O}$ (2)