

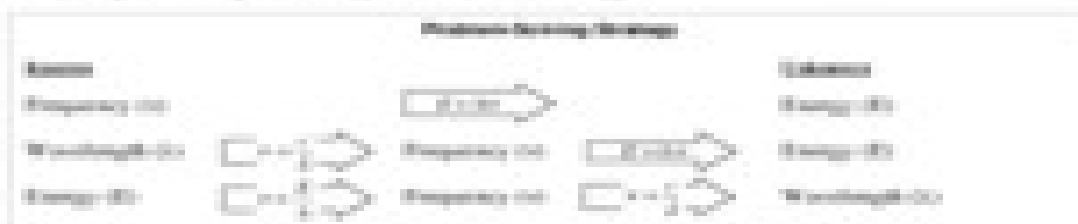
Planck's Equation

Chem Worksheet 5-2

Name _____

Max Planck showed that energy was transferred in discrete amounts called quanta, equal to $h\nu$. The variable h is a constant equal to 6.626×10^{-34} J·s and the variable ν represents the frequency in Hz. This equation allows us to calculate the energy of photons given their frequency. If the wavelength is given, the energy can be determined by first using the wave equation $c = \lambda\nu$ to calculate the frequency. Then using Planck's equation to calculate energy.

Unit Conversions	
1 m = 100 cm	1 cm = 1.00×10^{-2} m
1 nm = 10^{-9} m	1 m = 10^9 nm
1 $\mu\text{m} = 1 \times 10^{-6}$ m	1 nm = 10^{-9} m



Example

Light with a wavelength of 450 nm is given. Calculate the energy in joules of a given light photon.

- Given frequency: $\nu = 6.67 \times 10^{14}$ Hz
 - Wanted energy: $E = ?$ J
 - Planck's equation: $E = h\nu = (6.626 \times 10^{-34} \text{ J}\cdot\text{s})(6.67 \times 10^{14} \text{ Hz}) = 4.41 \times 10^{-19} \text{ J}$
 - Final energy: $E = 4.41 \times 10^{-19}$ J or 4.41×10^{-19} J photon

Use the equations above to answer the following questions.

- Electromagnetic radiation has a frequency of 6.61×10^{14} Hz. Calculate the energy, in joules, of the photon.
- Find the energy, in joules, for photons of electromagnetic radiation with a frequency of 1.01×10^{15} Hz.
- A certain type of lamp emits light photons with a wavelength of 1.00×10^{-7} m. What is the energy of these photons?
- One of the photons responsible for ultraviolet tanning produces ultraviolet light with a wavelength of 1.40×10^{-7} m. What amount of energy does the photon have?
- Find the energy in joules for a ray photon with a frequency of 3.4×10^{14} Hz.
- A red laser produces red light that has a wavelength of 700 nm. Calculate the energy in joules.
- What is the frequency of UV light that has an energy of 2.00×10^{-18} J?
- What is the wavelength and frequency of photons with an energy of 1.4×10^{-18} J?