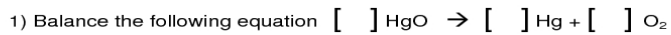


**Stoichiometry - Calculations**

Name: \_\_\_\_\_



2) The molar mass of **HgO** =

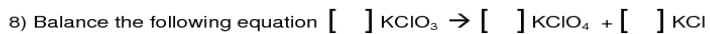
3) The molar mass of **O<sub>2</sub>** =

4) How many **moles of HgO** are there in **327.10 grams of HgO**?

5) How many **moles of O<sub>2</sub>** can be produced from **1.51 moles of HgO**?

6) How many **grams of O<sub>2</sub>** can be produced from **0.755 moles of HgO**?

7) How many **grams of O<sub>2</sub>** can be produced from **327.1 grams of HgO** during the decomposition reaction of HgO?



9) The molar mass of **KClO<sub>3</sub>** =

10) The molar mass of **KCl** =

11) How many **moles of KClO<sub>3</sub>** are there in **220.59 grams of KClO<sub>3</sub>**?

12) How many **moles of KCl** can be produced from **1.80 moles of KClO<sub>3</sub>**?

13) How many **grams of KCl** can be produced from **0.45 moles of KClO<sub>3</sub>**?

14) How many **grams of KCl** can be produced from **1.80 moles of KClO<sub>3</sub>** during the decomposition reaction of KClO<sub>3</sub>?