PRACTICAL WORK SHEET 6

GRADE 12

STANDARDISATION OF A HYDROCHLORIC ACID SOLUTION WITH A STANDARD SOLUTION OF SODIUM HYDROXIDE

INSTRUCTIONS

- You will work in groups.
- Follow the instructions as indicated under the procedure.
- Each learner must copy and complete the table.
- You will be assessed on your skills to measure accurately.

 - interpret data.
 - follow the procedures correctly.

To standardise a HCl - solution with the standard NaOH - solution.

EQUIPMENT

- -250ml Erlenmeyer flasks
- -25 ml pipette -burette
- -bromothymol blue -250ml reagent bottle
- -retort stand
- -wash-bottle
- -standard NaOH solution
- -mass meter
- -filter paper -500ml volumetric flask
- -funnel
- -10 ml measuring flask
- -concentrated HCI

PROCEDURE

- Prepare 500 ml of a ±0,1 mol.dm⁻³ HCl-solution.
 - Add 5 ml conc. HCl into the 500 ml volumetric flask. Fill the flask with water to the mark.

 - Shake well.
- Rinse the burette with a standard NaOH solution.
- Fill the burette with NaOH solution.
 Pipette 25 ml of the HCl-solution into each of three Erlenmeyer flasks.
- Add 2 3 drops BTB indicator (Bromothymol blue) to each flask.

 Titrate the acid (HCl) against the NaOH solution.

 The colour change will indicate the point of neutralisation (end point).
- Repeat the titration twice.

ASSESSMENT OF PROCEDURE

Criteria

- Correct apparatus selected.
- Correct acid selected.
- Correct base selected.
- Correct indicator selected.
- Wash down the inside of the volumetric flask with distilled water from time to time.
- Swirling of the mixture in the volumetric flask during the titration. Correct handling of burette and pipette. 6.
- Cleaning of the apparatus afterwards.