

Name _____ Period _____ Date _____

Chemistry Test – Chapters 8 & 9: Ionic and Covalent Bonds, Naming Compounds

- _____ 1. The reactivity of an element depends mainly on the number of:
a. valence electrons
b. inner electrons
c. energy levels
d. neutrons
- _____ 2. The force that holds two atoms together:
a. ionic bond
b. chemical bond
c. metallic bond
d. alloy
- _____ 3. A positively charged ion:
a. salt
b. alloy
c. cation
d. anion
- _____ 4. A negatively charged ion:
a. acid
b. alloy
c. anion
d. cation
- _____ 5. The electrostatic force that holds oppositely charged particles together:
a. ionic bond
b. metallic bond
c. lattice energy
d. covalent bond
- _____ 6. Where does a subscript that indicates the number of atoms appear, relative to a chemical symbol in a formula?
a. to the upper right
b. to the lower right
c. to the upper left
d. to the lower left
- _____ 7. A charged particle containing more than one atom:
a. cation
b. monatomic ion
c. polyatomic ion
d. alloy
- _____ 8. A charged particle contains only one atom:
a. salt
b. polyatomic ion
c. anion
d. monatomic ion
- _____ 9. The charge of monatomic ion is called its:
a. oxidation number
b. formula unit
c. chemical bond
d. lattice energy
- _____ 10. The particle formed when two or more atoms bond covalently:
a. salt
b. monomer
c. hybridization
d. molecule
- _____ 11. In the formation of a covalent bond, electrons are:
a. shared
b. gained
c. transferred
d. lost