

Name _____ Period _____ Date _____

Molar Conversions – Ch. 3 & 7

(p. 80-85, 221-226)

FOR ALL CALCULATIONS, SHOW YOUR WORK AND INCLUDE UNITS & SIG FIGS

PART A – MOLAR MASS

1. Calculate the molar mass for each of the following compounds. Include units!

calcium nitrate	Formula:	lead(II) iodide	Formula:
Molar mass:		Molar mass:	

PART B – MOLAR CONVERSIONS (Show your work and include units!)

2. How many moles of ammonia are in 1.20×10^{25} molecules of ammonia?

Formula:	Molar mass:
Answer:	

3. You need 2.5 moles of aluminum for an experiment. How many atoms of aluminum is this?

Formula:	Molar mass:
Answer:	