Name	answers	Date	<u> </u>
IONIC BOND Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.			
chemical family	ionic b	ond	octet rule
electron affinity	meta	al	shell
ion	nonnm	etal	valence electron
1. Each atom consists of a central nucleus and several shells that contain electrons. The outermost electronsare called valence electrons. (Inner electrons are not shown.) How many valence electrons does each atom have?			
sodium	1	chlorine	7
2. Click Pause (). Elements can be classified as metals and nonmetals. Metals do not hold on to theirvalence electrons very tightly, while nonmetals hold their electrons tightly. Electron affinity is a measure ofhow tightly the valence electrons are held.			
a. Try pulling an electron away from each atom. Based on this experiment, which atom is a metal? sodium which is non-metal chlorine			
b. Try moving an electron from the metal to the nonmetal. What happens?			
Sodium becomes a cation, and chlorine becomes an anion			
Introduction: Some of the particles that make up atoms have an electrical charge. Electrons are negativelycharged, while protons are positively charged. Particles with opposite charges (+ and -) attract, while particles with the same charge (+ and + or - and -) repel. Question: What happens when atoms gain or lose electrons? 1. Count: Electrons move around the nucleus of atoms in specific shells, shown by the rings around the atoms in the Gizmo. The first ring holds two electrons, and the second holds eight.			
(Electrons in the inner ringsare not shown; you can assume these rings are full.)			
A. Observe the sodium and chlorine atoms. Assuming that the inner rings are full of electrons, howmany electrons are there total in each atom?			
sodium 11	electrons	chlorine	17 electrons
B. Each atom is neutrally charged, which means that each atom has the same number of protons andelectrons. Based on this, how many protons are in each atom?			
sodium 1	1 protons	chlorine	17 prootns