

NAME:  
PERIOD:

**IB CHEMISTRY – CHAP. 7 FORMULA-NOMENCLATURE REVIEW WORKSHEET**

**General Rules for naming an ionic compound [metal ion or polyatomic ion bonded to a nonmetal ion or a different polyatomic ion]. Take the name of the ion off the ion list. If the metal ion has more than one charge listed on the ion list, use Roman numerals to indicate the charge.**

**General Rules for naming a molecular compound [nonmetal bonded to nonmetal]. The first name is the name of the first atom. If there is a subscript, use a prefix to indicate how many atoms. The second name is the name of the second atom, but change the name so it ends in -ide ex. Oxygen → oxide. Use a prefix to indicate how many.**

**DIRECTIONS: Part 1. Write the name for each of the compounds listed.**

CaO	_____	KBr	_____
NaCl	_____	Al(OH) <sub>3</sub>	_____
CaCl <sub>2</sub>	_____	PbSO <sub>4</sub>	_____
NH <sub>4</sub> NO <sub>3</sub>	_____	MgO	_____
Ca(OH) <sub>2</sub>	_____	LiF	_____
K <sub>2</sub> S	_____	Al <sub>2</sub> O <sub>3</sub>	_____
Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	_____	Na <sub>3</sub> PO <sub>4</sub>	_____
SiO <sub>2</sub>	_____	SO <sub>2</sub>	_____
SO <sub>3</sub>	_____	CF <sub>4</sub>	_____
PBr <sub>5</sub>	_____	P <sub>2</sub> O <sub>5</sub>	_____
N <sub>2</sub> O <sub>4</sub>	_____	PCl <sub>3</sub>	_____
N <sub>2</sub> O <sub>5</sub>	_____	N <sub>2</sub> O <sub>3</sub>	_____
CO	_____	PBr <sub>3</sub>	_____