ICE-CREAM LAB!

Purpose: The purpose of this experiment is to understand how much thermal energy is transfered from hot to cold substances.

Hypothesis: The hypothesis is that if smaller masses of ice-cream are compared, the thermal energy transfered will be less than the larger masses of ice-cream.

Backround: Calorimetry is measuring the amout of heat given out or taken in during a process such as combustion or change of state. The measurments are often made by observing the amount of solid liquefied, or liquid vaporizated, under set conditions. The Thermal energy transfered= 4.18J/(g)(K) x Mass (g) of cold water x Temperature change of cold water.

Important Observations: An Important Observation was that the hypothesis was correct. It was correct because the smaller the substance ~~you have~~, the amount of Thermal energy being tansfered will be less ~~Thermal energy~~; and ~~if you have~~ with more of the substance, ~~you will have~~ there will be more Thermal energy. So, this proves that the hypothesis is correct.

Some other Important Observations was that after the ice melted the heat that was taken away from the cream, made the cream freeze. But, the heat trensfered back into the cream once Team 4 wanted to make the ice-cream harder. This makes the heat and energy tranfer back into the cream because the ice was already melted. The ice can't take anymore heat in because the ice is already melted.

Another Important Observation was that the groups were mostly the same because they all had the same Spacific Heat, 4.18, and they all were close to the same temperature, 15 at the beginning and 0 at the end. But, they were different in ~~some ways because~~ that the masses were different.