Balancing Equations and Simple Stoichiometry-KEY

Balance the following equations:

1) 1 N_2 + 3 F_2 → 2 NF_3

2) 2 C_6H_{10} + 17 O_2 → 12 CO_2 + 10 H_2O

3) 1 HBr + 1 $KHCO_3$ → 1 H_2O + 1 KBr + 1 CO_2

4) 2 $GaBr_3$ + 3 Na_2SO_3 → 1 $Ga_2(SO_3)_3$ + 6 NaBr

5) 3 SnO + 2 NF_3 → 3 SnF_2 + 1 N_2O_3

Using the equation from problem 2 above, answer the following questions:

6) If I do this reaction with 35 grams of C_6H_{10} and 45 grams of oxygen, how many grams of carbon dioxide will be formed?
When you do this calculation for 35 grams of C_6H_{10}, you find that 113 grams of CO_2 will be formed. When you do the calculation for 45